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Chapter 10 Supplemental Problems Balance the following chemical equations. 1) $\text{SnS}_2(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{SnO}_2(\text{s}) + \text{SO}_2(\text{g})$ 2) $\text{C}_2\text{H}_6(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$ 3) $\text{Al}(\text{s}) + \text{HCl}(\text{aq}) \rightarrow \text{AlCl}_3(\text{aq}) + \text{H}_2(\text{g})$ 4) $\text{CoCO}_3(\text{s}) \rightarrow \text{CoO}(\text{s}) + \text{CO}_2(\text{g})$ Write a balanced equation for each of the following reactions, substituting symbols and formulas for names.

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SOLUTIONS MANUAL 3. Apply How many atoms of hydrogen-1 are in a 1.007-g sample? Recall that 1.007 amu is the mass of one atom of hydrogen-1. Round your answer to two significant digits. 1.007 g 1 H-1 atom 1
...

Chapter 10 Supplemental Problems Answer Key

Chapter 10 Supplemental Problems The Supplemental Problems for Chapter 10. Supplemental Problems for Chapter 10. 1) One overarching trend in atomic structures is that as the coordination number increases, the interatomic separation increases. Using the appropriate equations, answer the following questions. Supplemental Problems for Chapter 10

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