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**Chapter 16, Sections 10**

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*Structures The Compound:*

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Section 16.1 – The Nature of  
Covalent Bonding Practice

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Key Problems 1. Draw electron dot structures for each molecule. a. chlorine b. bromine c. iodine What do you observe about the three structures? They all look similar with respect to the electrons. 2. The following

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Key  
Molecules have single  
covalent bonds. Draw an  
electron dot structure for  
each. a.  $\text{H}_2\text{O}$  b.  $\text{PCl}_3$

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28, 30, 31, 35-37, 40,



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Key 43-46, Extra Credit: 50-53,  
55, 56, 58, 59, 62-67

Section 16.1 – The Nature of  
Covalent Bonding Practice  
Problems 1. Draw electron  
dot structures for each  
molecule. a. chlorine b.  
bromine c. iodine What do

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**Key** you observe about the three structures? 2. The following molecules have single covalent bonds.

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16 Problems Single Covalent Bonds 23. Describe the difference between an ionic and a covalent bond. 24. Classify the following compounds as ionic or covalent: a.  $MgCl_2$  b.  $Na_2S$  c.  $H_2O$  d.  $H_2S$  25.

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Key Explain why neon is monatomic but chlorine is diatomic. 26.

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Section 16.2 – Bonding  
Theories Section Review 16.2  
13. Use the molecular orbital theory to describe covalent bonding. What occurs during hybridization? When two atoms combine in a bonding situation, their



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Key  
atomic orbitals overlap to  
produce molecular orbitals.  
In hybridization, several  
atomic

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2. A  $\sigma$  bond is a covalent bond in which the electron density is concentrated in

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Key the region along the internuclear axis; that is, a line between the nuclei would pass through the center of the overlap region. CHAPTER 9 COVALENT BONDING: ORBITALS 3 6. acces pdf pearson education

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two elements in a covalent bond represent? answer choices . The process of electron gambling.



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