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Table OCR A 5.1.3

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as indicator | Acids

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Chapter 20 Acid And  
Bases

Chapter 20 acids and  
bases. the particle  
formed when a base  
gains a hydrogen ion:  
 $\text{NH}_4^+$  is the conjugate  
acid of the base  
 $\text{NH}_3$ \*20.3. two

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## Chapter 20 Acid

substances that are related by the loss or gain of a single hydrogen ion. Ammonia ( $\text{NH}_3$ ) and the ammonium ion ( $\text{NH}_4^+$ ) are a conjugate acid-base pair.

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AND BASES Section  
20.1 DESCRIBING  
ACIDS AND BASES  
Section 20.2

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Answers  
HYDROGEN IONS  
AND ACIDITY

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Baking Soda SMALL-

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cards. What is the

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Arrhenius definition of an acid and base?

Arrhenius acid increases  $H^+$  ions present when added to water.

Arrhenius base increases  $OH^-$  ions present when added to water. What is the hydronium ion?  $H_3O^+$  Formed when a hydrogen ion donated by an acid is bonded to a water molecule ...

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with Smolenyak ...

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Bases, and Redox. 20.1  
Acid and bases; 20.2  
Relative Strengths of  
Acids and Bases; 20.3  
Acidic, Basic, and  
Neutral Solutions; 20.A  
Acid Rain and Basic  
Oceans; 20.4 Losing  
and Gaining Electrons;

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## Chapter 20 Acid

20.5 Electrochemistry;  
20.6 Corrosion and  
Combustion

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### Chapter 20 Acids Bases Work Answers

1. Give three properties of acids and bases each. Acids are sour, turn litmus red, react with metals to produce hydrogen gas, react with bases to form salt and

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## Chapter 20 Acid

water, are corrosive.

Bases are bitter, turn litmus blue, feel slippery, react with acids to produce salt and water, are caustic.

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Bases. Acids. Base.

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Taste sour, turns litmus paper red, reacts with some metals to.... Tastes bitter, feels slippery, turns litmus paper blue, reacts.... compounds that form hydrogen ions when dissolved in water. A compound that produces hydroxide ions when dissolved in water.

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### And Bases

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casting of Page 3/8

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Bases, and Redox. 20.1

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*Page 25/32*

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Chapter 20 Acid  
Acids and Bases; 20.3  
Acidic, Basic, and  
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(Arrhenius) – produces  
*Page 28/32*

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## Chapter 20 Acid H<sup>+</sup>(aq) when dissolved in water.

### Answers

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Brønsted-Lowry Acids

& Bases Identify each

species in the following

equation as either the

Brønsted-Lowry acid,

the Brønsted-Lowry

base, the conjugate acid,

or the conjugate base.

Identify the conjugate

acid-base pairs in the

reaction.  $\text{H}_2\text{SO}_4$  (aq)

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## Chapter 20 Acid



Weak Acids and Bases

Strong ...

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