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mass relationships of elements in

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Stoichiometry. SECTION 2. PROBLEMS

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Write the answer on the line to the left.  
Show all your work in the space provided.

1. The following equation represents a laboratory preparation for oxygen gas:  
 $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  How many moles of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed?
2. Given the following equation:  $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$



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## *CHAPTER 9 REVIEW*

Chapter menu Resources Chapter 9

Section 1 Introduction to Stoichiometry

Objective • Define stoichiometry. •

Describe the importance of the mole ratio in stoichiometric calculations. • Write a mole ratio relating two substances in a

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chemical equation.

*Chapter 9 Stoichiometry Table of Contents*

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Terms in this set (12) Composition  
Stoichiometry. calculations involving the  
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compounds. Reaction Stoichiometry.  
calculations involving the mass  
relationships between reactants and  
products in a chemical reaction . Mole  
Ratio ...

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CHAPTER 9 REVIEW Stoichiometry

SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. 88%  
The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of  $H_2$  according to

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the following equation:  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$  ?  
2.0 mol a. Which chemical  
is ...

*mc06se cFMSr i-vi*

SECTION 1 Introduction to Stoichiometry

SECTION 2 Ideal Stoichiometric

Calculations SECTION 3 Limiting

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Reactants and Percentage Yield Why It Matters Video HMHScience.com GO ONLINE Stoichiometry BIG IDEA Reaction stoichiometry uses molar relationships to determine the amounts of unknown reactants or products from the amounts of known reactants or products. CHAPTER 9 DO NOT EDIT--Changes



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*CorrectionKey=NL-A DO NOT  
EDIT--Changes must be made ...*

Chapter 9. Solutions. Search for: 9.3  
Solution Stoichiometry. Learning  
Objectives. By the end of this section, you  
will be able to: Perform stoichiometric

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calculations involving solution molarity;  
As we have seen in lab, many reactions such as single or double displacement reactions are carried out in aqueous medium (i.e. in water). Because these reactions occur in aqueous solution, we can use ...

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## *9.3 Solution Stoichiometry / Introductory Chemistry*

Chapter 9 STOICHIOMETRY. What is Sto-Key-Om-It-Tree? Stoichiometry is derived from the Greek words stoicheion “element” and metron “measure”.

Stoichiometry is defined as the branch of chemistry that deals with the mass

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relationships between elements in a compound and the mass relationships between reactants and products in a chemical reaction. There are 2 main types of stoichiometry ...

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11.2. As you read this section, summarize each step on a tab and include an example of the step. 368 Chapter 11 •

Stoichiometry Section 11.11.1 Objectives

Describe the types of relationships indicated by a balanced chemical equation.

State the mole ratios from a balanced chemical equation. Review Vocabulary

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reactant: the starting substance in a ...

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