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ACIDS BASES &
SALTS-FULL

CHAPTER || CLASS 10
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IGCSE Chemistry: Acids
Bases and Salts pH,
pOH, H_3O^+ , OH^- , K_w ,
 K_a , K_b , pK_a , and pK_b

Basic Calculations -Acids
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 $[H_3O^+]$, $[OH^-]$ of
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Bases and Salts 01 :

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Acids Bases

Chemistry 12 Notes on
Unit 4 – Acids, Bases
and Salts Chemistry 12

-Unit 4 -Notes Page 2 -

Any acid (weak or
strong) could have high
or low concentration.

Weak & Strong → refers
to % ionization.

Concentration → the
moles of acid dissolved
per litre. Eg.) 10.0 M HCl

→ conc. and strong [H

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$3\text{O}^+] = 10.0 \text{ M } 0.001 \text{ M}$
HCl à dilute and strong
[H

Chem 12- Notes on
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Acids from HClO_4 to
 H_2SO_4 are 100% ionized
in water . only solvent
used in Chem 12 (and
most Chemistry) so even
though HClO_4 is above

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HCl on the chart, it is no more acidic in a water solution. H_3O^+ is the strongest acid that can exist in an undissociated form in water solution. all stronger acids ionize to form H_3O^+

Chem 12- Notes on
Acids & Bases - SSS
Chemistry

In this chapter, we will

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focus on acids and bases and their chemistry. 12.1: Introduction Certain household chemicals, such as some brands of cleanser, can be very concentrated bases, which makes them among the most potentially hazardous substances found around the home; if spilled on the skin, the strong caustic compound can

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immediately remove H^+ ions from the flesh, resulting in chemical burns.

12: Acids and Bases -
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Acids & Bases Chemistry
12 Notes on Unit 4 –
Acids, Bases and Salts
Chemistry 12 -Unit 4
-Notes Page 2 - Any acid

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(weak or strong) could have high or low concentration. Weak & Strong \rightarrow refers to % ionization.

Concentration \rightarrow the moles of acid dissolved per litre. Eg.) 10.0 M HCl \rightarrow conc. and strong $[H_3O^+] =$

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Unit 4 – Acids, Bases
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-Unit 4 -Notes Page 2 -

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$3\text{O}^+]$ = 10.0 M 0.001 M
HCl à dilute and strong
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Acids & Bases - SSS
Chemistry 12.4: Acid-
Base Titrations A
titration is the
quantitative reaction of
an acid and a base.

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Notes On Acids

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-Unit 4 -Notes Page 2 -

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à conc.

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Chemistry

Acids & Bases .

CHEMISTRY 12 UNIT

4 ACIDS, BASES AND

SALTS - OUTLINE

DETAILED NOTES ON

BRONSTED - LOWRY

THEORY (TUTORIAL

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14) Tutorial 14-Solutions
Arrhenius Theory
VIDEO. Br nsted-Lowry
Theory VIDEO.

DETAILED NOTES ON
STRONG AND WEAK
ACIDS AND ACID
BASE EQUILIBRIA
(Unit 4 p 1-12) Strong
Acids and Bases VIDEO.
Weak Acids VIDEO.
Weak Bases VIDEO

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Chemistry 12

Acids and Bases - Section
14 of General Chemistry

Notes is 23 pages in
length (page 14-1
through page 14-23) and
covers ALL you'll need
to know on the following
lecture/textbook topics:.

SECTION 14 - Acids
and Bases 14-1 -- Acid-
Base Chemistry ·

Arrhenius Acids and
Bases · Bronsted-Lowry

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Acids and Bases -
Proton (H^+) Donors
and Proton (H^+)
Acceptors

Chemistry Notes |
Chemistry Pdf -- Acids,
Bases, and the ...

1. The properties of acids
and bases 2. pH and
pOH calculations 3.
Definitions of acids and
bases 4. Neutralization

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reactions. NOTES:

Properties of acids : •
Form hydrogen ions
(H^+ or protons) in
solution • Form
hydronium (a water
molecule bonded to a
hydrogen ion shown as
 H_3O^+ or H^+).

CHEMISTRY NOTES
– CHAPTERS 20 AND
21 Acids and Bases ...

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Chemistry 12 Unit IV –
Acids, Bases and Salts

Notes IV.1 – The
Arrhenius Theory Of
Acids and Bases • Acid:
any substance which
releases H^+ (aq) in
water. • Base: any

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substance which releases
 OH^- (aq) in water. •

Salt: is the neutralization
product which results
when an acid and a base
react: HCl (aq) + NaOH
 $(\text{aq}) \rightarrow \text{NaCl (aq) + H}_2\text{O}$
(l ...

Chemistry 12

Acids are classified as
Organic Acids and
Mineral Acids. Acids

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Acids and Bases
SSS Chemistry

which are derived from plants and animals, they are known as Organic Acids. For Example, Citric Acid from fruit.

Mineral acids are inorganic acids such as Sulphuric Acid. They are dangerous to be used, so need more precautions. Acids are also classified as Strong Acids or Weak Acids. Strong acid is an acid, that completely

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dissociates into ions in
aqueous solutions.

Revision Notes for
Science Chapter 2 -
Acids, bases and ...
Acid-Base Titrations.
Indicators. Titration
Curves & Solving AB
Titration Problems
(Hebden summary) Acid
Rain. Notes from demo:
Lab Instructions Acid-

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Base Titration Suggested
Questions to prep for
written component of
Acid/Base Unit Exam:
Hebden Q #: 67, 69.e),
99, 101, 104, 113, 115,
118, 120, 125.e), 126,
130, 139, 140

Unit 4 Acids, Bases, and
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Acids and Bases, Buffer
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Curves, pKa and pH)

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Chapter 12 Notes ...
Phosphoric acid (H_3PO_4) etc. Chemical
Properties of Acid: (i)

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Reaction of acids with metal: Acids give hydrogen gas along with respective salt when they react with a metal. Metal + Acid → Salt +

Hydrogen Examples:
Hydrogen gas and zinc chloride are formed when hydrochloric acid reacts with zinc metal.

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Class Notes: I) Weak

Acid Equilibrium and

K_a; II) Weak Base

Equilibrium and K_b; III)

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Writing Formula
(Molecular), Complete
Ionic and Net Ionic
Equations for Acid/Base

...

Chemistry 12 - Miss
Zukowski's Class
Notes Chapter - 2 Acids,
Bases and Salts ACIDS:

- These are the substances which have sour in taste.
- They

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turns blue litmus solution to red. • They give H^+ ions in aqueous solution

Examples of Acids : -

HCl - Hydrochloric

Acid BASES • These

substances are bitter in

taste. • They turns red

litmus solution to blue.

Chapter - 2, Acids Bases
and Salts Notes

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