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Discussion In Experiment 13, we compared the freezing point of stearic acid with that of solutions of stearic acid and an increasing amount of unknown fatty acid in order to find the molar mass and identity of the unknown. The first part of the lab measured the freezing point of stearic acid with no unknown. The graph produced from the data,

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(Experiment 13.1) $\text{HCl (aq)} + \text{H}_2\text{O (l)} \rightarrow \text{H}_3\text{O}^+ \text{ (aq)} + \text{Cl}^- \text{ (aq)}$ (Experiment 13.2) $\text{HCl (aq)} \rightarrow \text{H}^+ \text{ (aq)} + \text{Cl}^- \text{ (aq)}$ Strong bases are substances that ionize completely in water producing hydroxide ions. The ionization of sodium hydroxide, a strong base, is shown as the simplified form in Experiment 13.3.

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As the world of chemistry constantly evolves, certain elements, compounds, and even substances are always undergoing substance identification. Moreover, there are many methods to identify a substance. As noted by Austin

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