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DH . Work is found the first law as $w = q - \Delta U$; thus $q = DH$; $w = -DHPVL$; 4. Isothermal Process
Because U is a function only of T for an ideal gas, $\Delta U = \Delta H = 0$ for an isothermal process. These results also follow from the general results by using $D\ln(TV) = 0$ for an isothermal process.

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