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1 point for answer) in 100.0 g sample 1.00 g
/ mol 85.7 g 12.0 g/ mol 14.3 mol H 7.14 mol
C - 6.0 6 (CH₂) 14.3 85.7% c 84.0 u = 14.3

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mol H = 7.14 mol C 2 = CH₂ empirical formula
6 12 (1 point) (7 point for method, 1 point
for answer for each part) 12+2 u 10.

molecules x - - 8 molecules H₂O - 9.0 mol CO₂
3.0 mol - -x + 18 g. 11 g +40. g -

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